

21. Code B
22. The energy level for sulfur "S" is
a) s b) p c) 2 d) 3
23. How many orbitals are present in the 'd' subshell?
a) 1 b) 3 c) 5 d) 6 e) 10
24. The sodium ion Na^+ has the same electronic configuration as
a) Ar b) Ne c) Mg d) Kr
25. The maximum number of electrons in an orbital is _____.
a) 1 b) 2 c) 6 d) 14
26. The halogens end in _____
a) s^2 b) p^6 c) p^7 d) d^3 e) p^5
27. Which element ends in $4p^3$?
a) As b) P c) N d) S e) Se
28. How many electrons can the subshell 'p' hold?
a) 3 b) 5 c) 6 d) 7 e) 8
29. If an element ends in p^4 , how many electrons are paired?
a) all are paired b) none are paired c) 2 are paired
30. Which of the following has a different electronic structure from the others?
a) Ar b) O^{-2} c) F^{-1} d) Na^{+1}