

81. Code A

82. How many grams are present in 0.33 moles of Al_2O_3 ?

- a) 67.3 b) 33.7 c) 0.333 d) 3.37

83. How many moles of CaCl_2 are present in 138.8 g of CaCl_2 ?

- a) 1.25 b) 2.50 c) 0.063 d) 128.8

84. How many **molecules** are present in 4 grams of CH_4 ?

- a) 2 b) 0.25 c) 1.51×10^{23} d) 2.41×10^{23}

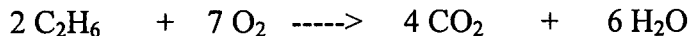
85. Given 1.51×10^{23} molecules of H_2O . How many moles of water are present?

- a) 0.25 b) 4.00 c) 18 d) 4.5

86. Given that 0.5 moles of a compound has a mass of 22 grams. What is the compound?

- a) O_2 b) N_2 c) CO_2 d) NH_3

87 - 90 Given



87. What is the molecular weight of C_2H_6 ?

- a) 30 b) 45 c) 26 d) 36

88. Given 1.5 grams of C_2H_6 , how many grams of CO_2 would be generated?

- a) 1.1 b) 2.2 c) 4.4 d) 8.8

89. How many grams of H_2O would be generated for each mole of C_2H_6 ?

- a) 8 b) 54 c) 27 d) 36

90. How many moles of CO_2 would be generated for each mole of C_2H_6 ?

- a) 1 b) 2 c) 3 d) 4 e) 5