41. Code A

42. What is the most abundant element in the universe?
a) H b) He c) Fe d) O

43. Which element has reactions similar to Oxygen?
a) Al b) N c) Cl d) Se

44. Isotopes have
a) the same atomic number and the same atomic mass
b) different atomic number and different atomic mass
c) same atomic number and different atomic mass
d) same atomic mass and different atomic number

45. What is the energy level of S (sulfur)?
a) 1 b) 2 c) 3 d) 4 e) 5

46. Which element ends in the following electronic notation $2p^1$?
a) Al b) B c) C d) O

47. Given $4p^2$, these electrons are
a) paired b) unpaired

48. Which of the following does **NOT** the same electronic formula as the others?

a) Kr b) Br^{-1} c) Ca^{+2} d) Rb^{+1} e) Sr^{+2}

---

49 - 51 Given

$$^{138}_{56}X^{+2}$$

49. How many electrons are represented here?
a) 54 b) 55 c) 56 d) 138 e) 2

50. How many neutrons does it have?
a) 138 b) 56 c) 54 d) 82

51. This is a(n) _____
a) anion b) bunion c) cation

---

52. The density of a substance is 4.6 g/mL and had a volume of 72 mL. What is its mass?
a) 331 g b) 15.7 g c) 0.64 g

53. A nickel has a mass of 5.0 grams expressed in nanograms is 5.9×10^{X} What is the value for x?
a) 3 b) -3 c) 9 d) -9 e) 12